

Centre Number						Candidate Number					
Surname						Other Names					
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Candidate Declaration. I have read and understood the Notice to Candidate and can confirm that I have produced the attached work without assistance other than that which is acceptable under the scheme of assessment.											
Candidate Signature						Date					

For Teacher's Use	
Section	Mark
PSA	
Task	
Section A	
Section B	
TOTAL (max 50)	



General Certificate of Education
Advanced Level Examination
June 2011

Chemistry

CHM6T/Q11/test

Unit 6T A2 Investigative Skills Assignment

For submission by 15 May 2011

For this paper you must have: <ul style="list-style-type: none"> the Periodic Table/Data Sheet provided at the end of this paper the Task Sheet and your Candidate Results Sheet a ruler with millimetre measurements a calculator. 	Time allowed <ul style="list-style-type: none"> 1 hour
Instructions: <ul style="list-style-type: none"> Use black ink or black ball-point pen. Fill in the boxes at the top of this page. Answer all questions. You must answer the questions in the space provided. Do not write outside the box around each page or on blank pages. Do all rough work in this book. Cross through any work you do not want to be marked. 	Information <ul style="list-style-type: none"> The marks for questions are shown in brackets. The maximum mark for this paper is 30. You will be marked on your ability to: <ul style="list-style-type: none"> organise information clearly use scientific terminology accurately.
Details of additional assistance (if any). Did the candidate receive any help or information in the production of this work? If you answer yes give the details below or on a separate page. Yes <input type="checkbox"/> No <input type="checkbox"/>	

Teacher Declaration:

I confirm that the candidate's work was conducted under the conditions laid out by the specification. I have authenticated the candidate's work and am satisfied that to the best of my knowledge the work produced is solely that of the candidate.

Signature of teacher Date

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Section A

These questions are about the task, some reactions of iron(II) and iron(III) compounds.

You should use your Task Sheet and your Candidate Results Sheet to answer them.

Answer **all** questions in the spaces provided.

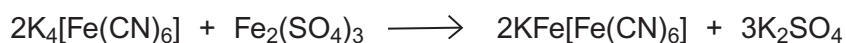
- 1** *Prussian Blue* dye is an insoluble complex with the formula $\text{KFe}[\text{Fe}(\text{CN})_6]$. Calculate the M_r of this complex using data from the Periodic Table. Give your answer to one decimal place.

.....
.....
(1 mark)

- 2** In one investigation a student obtained 1.65 g of *Prussian Blue*. Calculate the amount, in moles, of *Prussian Blue* obtained in this investigation.

.....
.....
(1 mark)

- 3** The iron(III) salt used in the Task was in excess. The 10.0 cm^3 solution of potassium hexacyanoferrate(II) had a concentration of 0.75 mol dm^{-3} . The equation for the formation of *Prussian Blue* from iron(III) sulfate and potassium hexacyanoferrate(II) is shown below.



Calculate the maximum amount, in moles, of *Prussian Blue* that could have been formed in the reaction.
Show your working.

.....
.....
.....
(2 marks)

- 4** Use your answers to Question 2 and Question 3 to calculate the percentage yield of *Prussian Blue* obtained in this experiment.

.....
.....
(1 mark)

- 5 *Prussian Blue* dye is an insoluble solid and when prepared as described in **Part 1** it is impure.
Suggest how the impurities present in the dye could be removed.
-
-
- (1 mark)
- 6 Give the formula for each of the iron-containing products formed in **Test 2**.
- Product using iron(II) sulfate solution
- Product using iron(III) nitrate solution
- (2 marks)
- 7 Write an ionic equation for the reaction of aqueous $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$ ions with hydroxide ions to form a brown precipitate.
-
- (1 mark)

9

Turn over for the next question

Turn over ►

Section B

Answer **all** questions in the spaces provided.

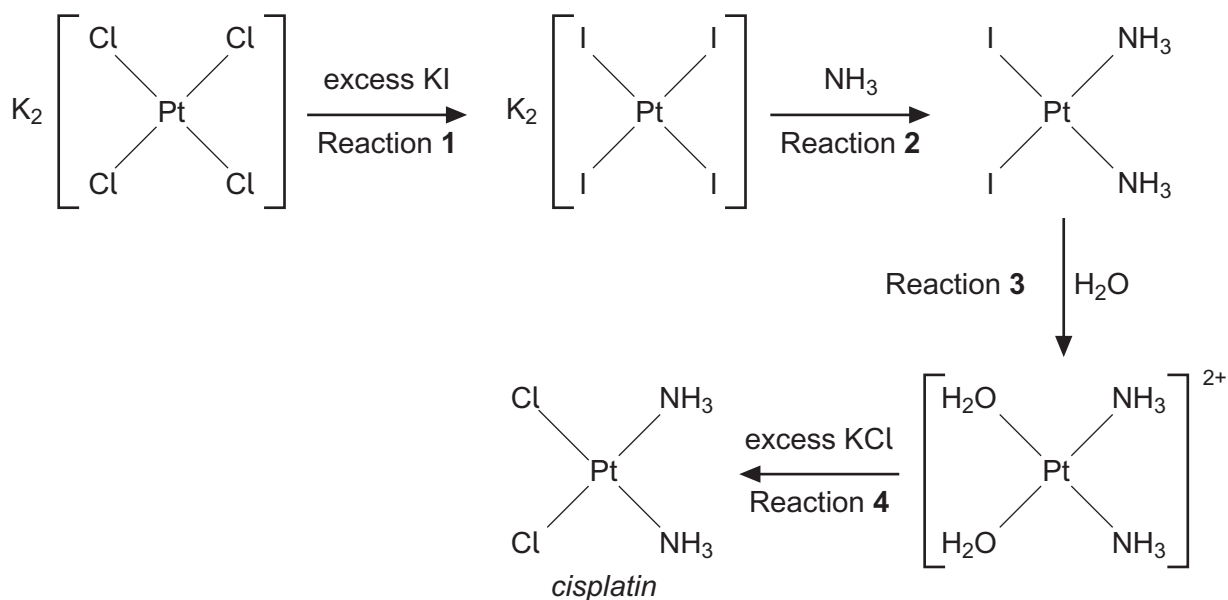
Introduction

Complexes containing transition elements have a wide variety of uses including acting as dyestuffs like *Prussian Blue*.

Cisplatin is a platinum-based chemotherapy drug used to treat various types of cancers. It was the first member of a class of anti-cancer drugs that react with DNA in tumour cells.

Cisplatin is prepared from K_2PtCl_4 according to the following scheme.

All the reactions shown are reversible.



- 8 Name the type of reaction occurring in all four steps of the scheme.

.....
(1 mark)

- 9 Explain why an excess of potassium iodide is used in Reaction 1.

.....
.....
.....
(2 marks)

10 (a) Write an equation for Reaction 1.

.....
.....
(1 mark)

10 (b) Calculate the percentage atom economy for the formation of K_2PtCl_4 in Reaction 1.
Show your working.

.....
.....
.....
.....
(2 marks)

11 In Reaction 3, silver nitrate solution is added to improve the yield of product.

11 (a) Write the **simplest ionic** equation for the reaction of iodide ions with silver nitrate.

.....
(1 mark)

11 (b) Suggest why addition of silver nitrate improves the yield of product from Reaction 3.

.....
.....
(1 mark)

12 Suggest two reasons, other than poor practical technique, why the overall yield of *cisplatin* in this synthesis may be low.

Reason 1

.....

Reason 2

.....

(2 marks)

Turn over ►

13 The *cisplatin* formed in Reaction 4 is impure. Outline how the impure solid is purified by recrystallisation.

.....

.....

.....

.....

.....

(3 marks)

14 Platinum compounds are highly toxic.

14 (a) State why *cisplatin* is used in cancer treatment despite its toxicity.

.....

.....

(1 mark)

14 (b) Suggest a suitable precaution that should be taken by medical staff when using *cisplatin*.

.....

(1 mark)

15 Other metal ions are also toxic to humans. A substance that can be used to treat such poisoning contains the ion EDTA^{4-} . EDTA^{4-} forms very stable complexes with metal ions. These complexes are **not** toxic.

15 (a) Write an equation for the reaction of EDTA^{4-} with aqueous copper(II) ions, $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$

.....

(1 mark)

- 15 (b)** A solution containing EDTA^{4-} can also be used in a titration to determine the concentration of metal ions in solution.
A river was polluted with copper(II) ions. When a 25.0 cm^3 sample of the river water was titrated with a $0.0150 \text{ mol dm}^{-3}$ solution of EDTA^{4-} , 6.45 cm^3 were required for complete reaction.
Calculate the concentration, in mol dm^{-3} , of copper(II) ions in the river water.
Show your working.
-
-
-
-
- (2 marks)

- 16** The determination of the concentration of copper(II) ions in a single sample of river water gives an unreliable value for the copper(II) ion pollution in the river.
Give **one** reason why this value is unreliable.
-
-
- (1 mark)

- 17** Silver complexes can be used to identify a particular organic functional group.
Give **one** example of a silver complex that can be used in this way and state the organic functional group it identifies.
- Silver complex
- Organic functional group
- (2 marks)

END OF QUESTIONS

GCE Chemistry Data Sheet

Table 1

Infrared absorption data

Bond	Wavenumber /cm ⁻¹
N-H (amines)	3300 – 3500
O-H (alcohols)	3230 – 3550
C-H	2850 – 3300
O-H (acids)	2500 – 3000
C≡N	2220 – 2260
C=O	1680 – 1750
C=C	1620 – 1680
C-O	1000 – 1300
C-C	750 – 1100

Table 2

¹H n.m.r. chemical shift data

Type of proton	δ/ppm
ROH	0.5–5.0
RCH ₃	0.7–1.2
RNH ₂	1.0–4.5
R ₂ CH ₂	1.2–1.4
R ₃ CH	1.4–1.6
	2.1–2.6
	3.1–3.9
RCH ₂ Cl or Br	3.1–4.2
	3.7–4.1
	4.5–6.0
	9.0–10.0
	10.0–12.0

Table 3

¹³C n.m.r. chemical shift data

Type of carbon	δ/ppm
	5–40
	10–70
	20–50
	25–60
	50–90
	90–150
	110–125
	110–160
	160–185
	190–220

