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Resource provided as an example of good practice by Ashley Law, Yarm School, Yarm.

Supported by



Research Notes

3. List below the sources which you used to complete this task.

1.....
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2.....
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3.....
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4. Explain why you think that each of the sources you listed above are suitable

Source 1
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Source 2
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Source 3
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Go through the appropriate sections of LabSkills on the Learning Platform, and then answer the following questions.

1. What do you need to make sure that you do when you set up your burette and clamp stand?

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2. What do you need to use to rinse the burette?

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3. What is the correct height to have the burette when filling it?

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4. What must be done when you have filled the burette with the reagent?

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5. How do you remove any air bubbles from the burette?

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6. The second reagent is then usually added to the conical flask. What is used to do this? What precautions do you need to take?

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7. What do you need to do with the conical flask when you are performing the titration? Suggest why this is important.

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8. Why is it helpful to use a white tile?

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9. When you are close to the end point, what do you do to make the end point more accurate?

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Titration Safety

10. What must you remember not to do when filling a burette?

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11. Why would you need to make sure that the burette is not clamped too loosely?

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12. Why would you need to make sure that the burette is not clamped too tightly?

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13. At what point on the burette should you clamp it?

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14. State what you use each of your hands for during the titration.

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Look at the section on "acid-base titration"

On the next page, write down the 6 stages involved in the acid-base titration.

Acid-Base Titration

6 stages involved are:

1	
2	
3	
4	
5	
6	