



Surname _____

Forename(s) _____

Centre Number _____

Candidate Number _____

Candidate Signature _____

I declare this is my own work.

GCSE

CHEMISTRY

H

Higher Tier Paper 1

8462/1H

Monday 22 May 2023

Morning

Time allowed: 1 hour 45 minutes

At the top of the page, write your surname and forename(s), your centre number, your candidate number and add your signature.

[Turn over]



MATERIALS

For this paper you must have:

- a ruler
- a scientific calculator
- the periodic table (enclosed).

INSTRUCTIONS

- Use black ink or black ball-point pen.
- Pencil should only be used for drawing.
- Answer ALL questions in the spaces provided.
Do not write on blank pages.
- If you need extra space for your answer(s), use the lined pages at the end of this book. Write the question number against your answer(s).
- Do all rough work in this book. Cross through any work you do not want to be marked.
- In all calculations, show clearly how you work out your answer.



INFORMATION

- **The maximum mark for this paper is 100.**
- **The marks for questions are shown in brackets.**
- **You are expected to use a calculator where appropriate.**
- **You are reminded of the need for good English and clear presentation in your answers.**

DO NOT TURN OVER UNTIL TOLD TO DO SO



0	1
---	---

Discoveries in chemistry led to a better understanding of atomic structure.

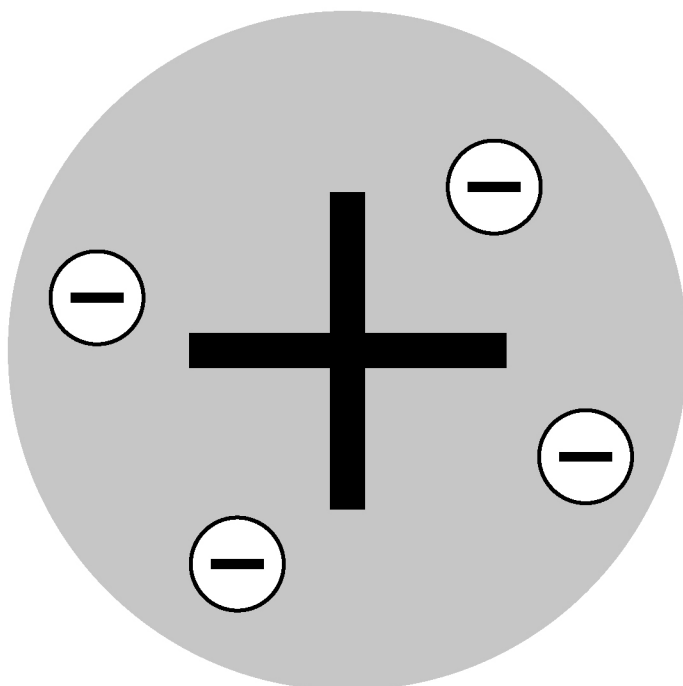
0	1	.	1
---	---	---	---

Atoms were originally thought to be tiny spheres that could not be divided.

The plum pudding model of the atom was then developed.

FIGURE 1 represents the plum pudding model of the atom.

FIGURE 1



**Describe the plum pudding model of the atom.
[2 marks]**

0 1 . 2

Atoms contain electrons, neutrons and protons.

**Write these three particles in order of their discovery.
[1 mark]**

Earliest _____

Latest _____

[Turn over]



Very few atoms of the element tennessine (Ts) have ever been identified.

The atomic number of tennessine is 117

0 1 . 3

Predict the number of outer shell electrons in an atom of tennessine.

Give ONE reason for your answer.

Use the periodic table. [2 marks]

Number of outer shell electrons _____

Reason _____



0	1	.	4
---	---	---	---

Tennessine was first identified by a small group of scientists in 2010.

Suggest ONE reason why tennessine was NOT accepted as a new element by other scientists until 2015. [1 mark]

[Turn over]



0	1	.	5
---	---	---	---

The discovery of isotopes explained why some relative atomic masses are not whole numbers.

Element R has two isotopes.

TABLE 1 shows the mass numbers and percentage abundances of the isotopes of element R.

TABLE 1

MASS NUMBER	PERCENTAGE ABUNDANCE (%)
6	7.6
7	92.4



Calculate the relative atomic mass (A_r) of element R.

Give your answer to 1 decimal place. [3 marks]

Relative atomic mass (1 decimal place) = _____

[Turn over]

9



0	2
---	---

This question is about temperature changes.

A student investigated the change in temperature of a solution when different masses of ammonium nitrate were dissolved in water.

This is the method used.

- 1. Measure 200 cm³ of water into a polystyrene cup.**
- 2. Measure the temperature of the water.**
- 3. Add 4.0 g of ammonium nitrate to the water.**
- 4. Stir the solution until all the ammonium nitrate has dissolved.**
- 5. Measure the lowest temperature reached by the solution.**
- 6. Repeat steps 1 to 5 with different masses of ammonium nitrate.**



0	2	.	1
---	---	---	---

Give the independent variable and the dependent variable in the investigation. [2 marks]

Independent variable _____

Dependent variable _____

[Turn over]



TABLE 2 shows the results.

TABLE 2

Mass of ammonium nitrate added in grams	Lowest temperature of solution in °C
4.0	18.2
8.0	16.2
12.0	15.2
16.0	13.6
20.0	12.4
24.0	10.6

0 2 . 2

Plot the data from TABLE 2 on FIGURE 2, on the opposite page.

Draw a line of best fit. [3 marks]

0 2 . 3

Determine the initial temperature of the water.

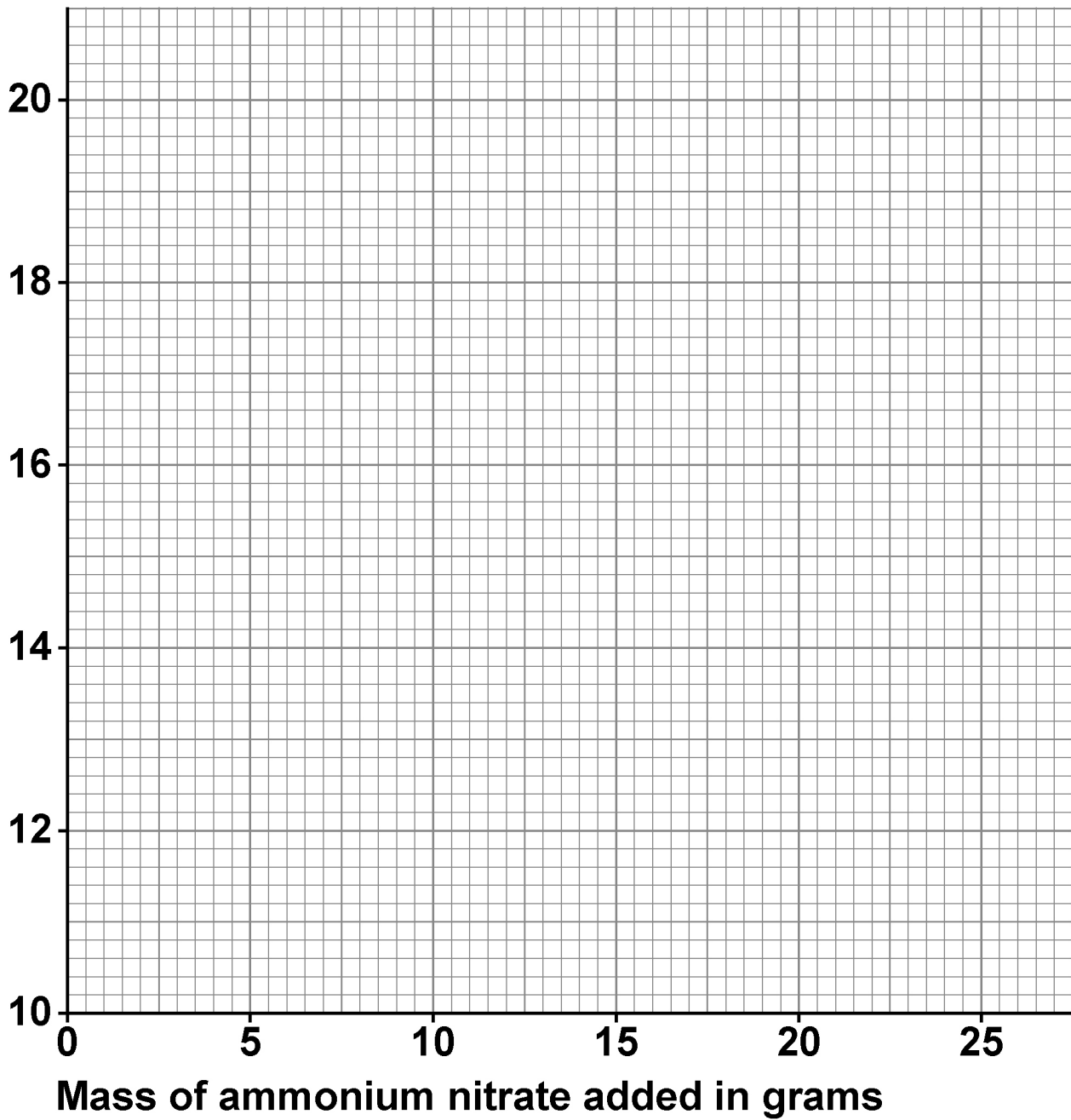
You should extend your line of best fit on FIGURE 2.
[2 marks]

Initial temperature of the water = _____ °C



FIGURE 2

Lowest temperature of solution in °C



[Turn over]



0	2	.	4
---	---	---	---

How do the results show that dissolving ammonium nitrate in water is endothermic? [1 mark]



BLANK PAGE

[Turn over]



The student repeated the experiment three more times.

TABLE 3 shows the results for 8.0 g of ammonium nitrate.

TABLE 3

	Trial 1	Trial 2	Trial 3	Trial 4	Mean
Lowest temperature of solution in °C	16.2	16.6	16.8	16.4	16.5

02.5

The student recorded the mean lowest temperature of the solution for 8.0 g of ammonium nitrate as 16.5 ± 0.3 °C.

Explain why the student included ± 0.3 °C after the mean lowest temperature. [2 marks]



0	2	.	6
---	---	---	---

What type of error is shown by the results in TABLE 3?
[1 mark]

Tick (✓) ONE box.

☐

Random error

☐

Systematic error

☐

Zero error

[Turn over]

<hr/>
11



0	3
---	---

This question is about making a soluble salt.

0	3	.	1
---	---	---	---

Plan a method to make pure, dry crystals of zinc chloride from zinc carbonate and a dilute acid.

[6 marks]



0	3	.	2
---	---	---	---

Name TWO other substances that can each be reacted with a dilute acid to make zinc chloride.

Do NOT refer to zinc carbonate in your answer.
[2 marks]

1 _____

2 _____

—
8



BLANK PAGE

[Turn over]



04

This question is about hydrogen and compounds of hydrogen.

FIGURE 3 shows the displayed formulae for the reaction between hydrogen and chlorine.

FIGURE 3



TABLE 4 shows the bond energies.

TABLE 4

Bond	H — H	Cl — Cl	H — Cl
Bond energy in kJ/mol	436	346	432



0	4	.	1
---	---	---	---

Which expression shows how to calculate the overall energy change for the reaction in FIGURE 3?

Use TABLE 4. [1 mark]

Tick (✓) ONE box.

☐

$$436 + 346 + 432 \text{ kJ/mol}$$

☐

$$436 + 346 + (2 \times 432) \text{ kJ/mol}$$

☐

$$436 + 346 - 432 \text{ kJ/mol}$$

☐

$$436 + 346 - (2 \times 432) \text{ kJ/mol}$$

[Turn over]



The reaction between hydrogen and chlorine is exothermic.

0 4 . 2

Explain why this reaction releases energy to the surroundings. [2 marks]

04.3

FIGURE 4 shows part of a reaction profile for the reaction between hydrogen and chlorine.

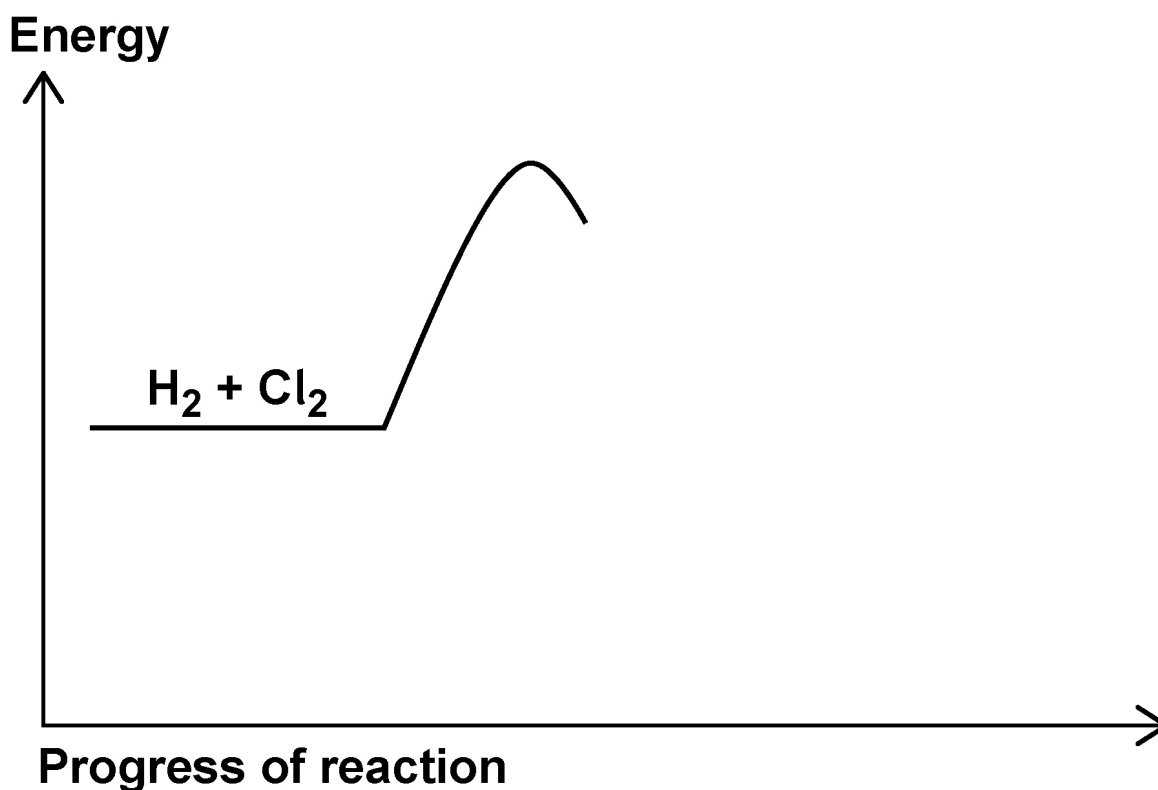
Complete the reaction profile in FIGURE 4.

You should:

- label the activation energy
- label the overall energy change.

[3 marks]

FIGURE 4



[Turn over]



0	4	.	4
---	---	---	---

Draw a dot and cross diagram for a molecule of hydrogen chloride (HCl).

Show the outer shell electrons only. [2 marks]

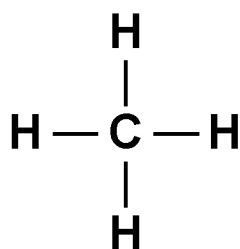


04.5

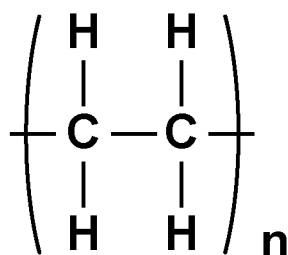
FIGURE 5 represents molecules of methane and of poly(ethene).

FIGURE 5

Methane



Poly(ethene)



Methane is a gas at room temperature but poly(ethene) is a solid at room temperature.

Explain why methane and poly(ethene) exist in different states at room temperature. [4 marks]

[Turn over]



12



0	5
---	---

This question is about acids and alkalis.

0	5	.	1
---	---	---	---

Ethanoic acid is a weak acid.

What is meant by 'weak acid'?

Answer in terms of ionisation. [1 mark]

[Turn over]



0	5	.	2
---	---	---	---

The concentration of an acid can be measured in mol/dm^3 .

Which combination of changes **INCREASES** the concentration of an acid? [1 mark]

Tick (✓) **ONE** box.

☐

The mass of acid dissolved is halved and the volume of the solution is halved.

☐

The mass of acid dissolved is halved and the volume of the solution is doubled.

☐

The mass of acid dissolved is doubled and the volume of the solution is halved.

☐

The mass of acid dissolved is doubled and the volume of the solution is doubled.



0	5	.	3
---	---	---	---

The concentration of an acid can be determined by titration.

An indicator is added to an alkali in a flask.

Name an indicator that can be used in this titration.

Give the colour change of the indicator when acid from a burette is added to the alkali in the flask. [2 marks]

Name of indicator _____

Colour change from _____ to _____

0	5	.	4
---	---	---	---

Sodium carbonate dissolves in water to produce an alkaline solution.

Give the formula of the ion that makes a solution alkaline. [1 mark]

[Turn over]



0	5	.	5
---	---	---	---

A student does a titration using sodium carbonate solution and nitric acid.

The equation for the reaction is:



25.0 cm³ of 0.124 mol/dm³ sodium carbonate solution is neutralised by 23.6 cm³ of nitric acid.

Calculate the concentration of the nitric acid.

Give your answer to 3 significant figures.

You should calculate:

- the number of moles of sodium carbonate in 25.0 cm³ of the solution
- the number of moles of nitric acid in 23.6 cm³ of the nitric acid
- the concentration of the nitric acid in mol/dm³.

[5 marks]



_____ mol/dm³

When hydrochloric acid dissolves in water, hydrogen ions (H^+) and chloride ions (Cl^-) are produced.

05.6

A solution of hydrochloric acid with pH 4.5 has a concentration of H^+ ions of $3.16 \times 10^{-5} \text{ mol/dm}^3$.

What is the concentration of H^+ ions in a solution of hydrochloric acid with pH 2.5? [1 mark]

Concentration of H^+ ions = _____ mol/dm^3

05.7

Which element has atoms that have the same electronic structure as the chloride ion?

Use the periodic table. [1 mark]

12



BLANK PAGE

[Turn over]



0	6
---	---

This question is about uses of metals in electrical wires.

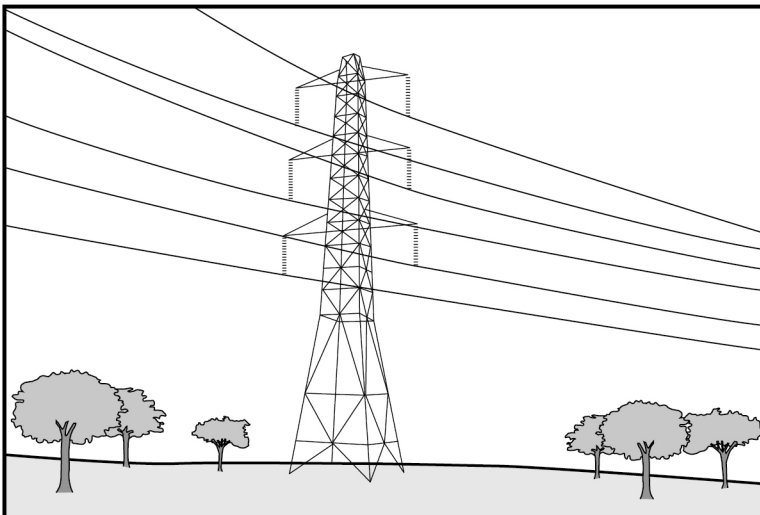
Electrical wires can be made from:

- aluminium
- copper
- silver.

FIGURE 6, on the opposite page, shows three uses of electrical wires.

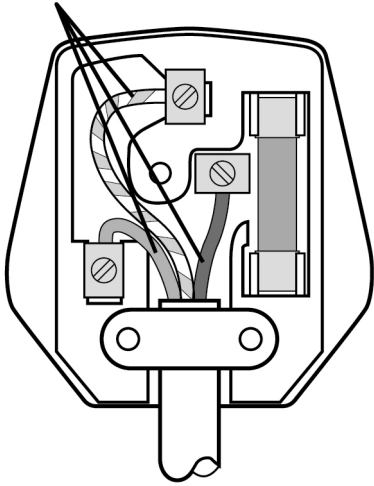
FIGURE 6

Overhead power cables

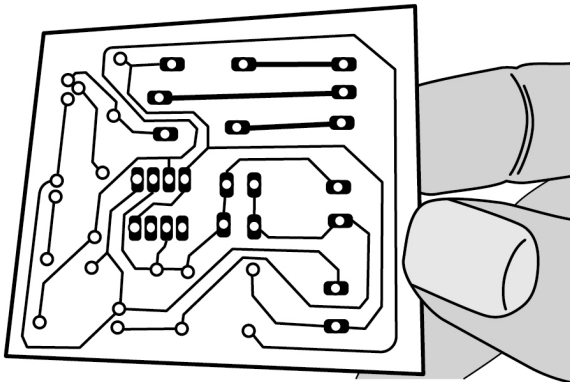


Wiring in homes

Electrical wires



Printed circuit boards



[Turn over]



TABLE 5 shows information about the metals.

The higher the value for electrical conductivity, the better the metal is at conducting electricity.

TABLE 5

	ALUMINIUM	COPPER	SILVER
Electrical conductivity in arbitrary units	37.7	59.6	63.0
Density in g/cm³	2.7	9.0	10.5
Cost of metal per kg in £	1.50	7.00	640.00

06.1

Evaluate the use of aluminium, copper and silver for the types of electrical wires shown in FIGURE 6, on page 37.

Use TABLE 5. [4 marks]



[illegible]

[Turn over]



0	6	.	2
---	---	---	---

Describe how metals conduct electricity. [3 marks]



0	6	.	3
---	---	---	---

Electrical wires are usually made of pure metals and NOT alloys. This is because pure metals are better electrical conductors.

Suggest why alloys do NOT conduct electricity as well as pure metals.

Answer in terms of structure and bonding. [2 marks]

[Turn over]

—
9



0	7
---	---

This question is about electrolysis.

Aluminium is manufactured by electrolysis of a molten mixture of aluminium oxide (Al_2O_3) and cryolite (Na_3AlF_6).

0	7	.	1
---	---	---	---

Complete the half equation for the reaction occurring at the negative electrode. [1 mark]



0	7	.	2
---	---	---	---

Cryolite contains Na^{+} ions as well as Al^{3+} ions.

Suggest ONE reason why sodium is NOT a product of the electrolysis. [1 mark]



BLANK PAGE

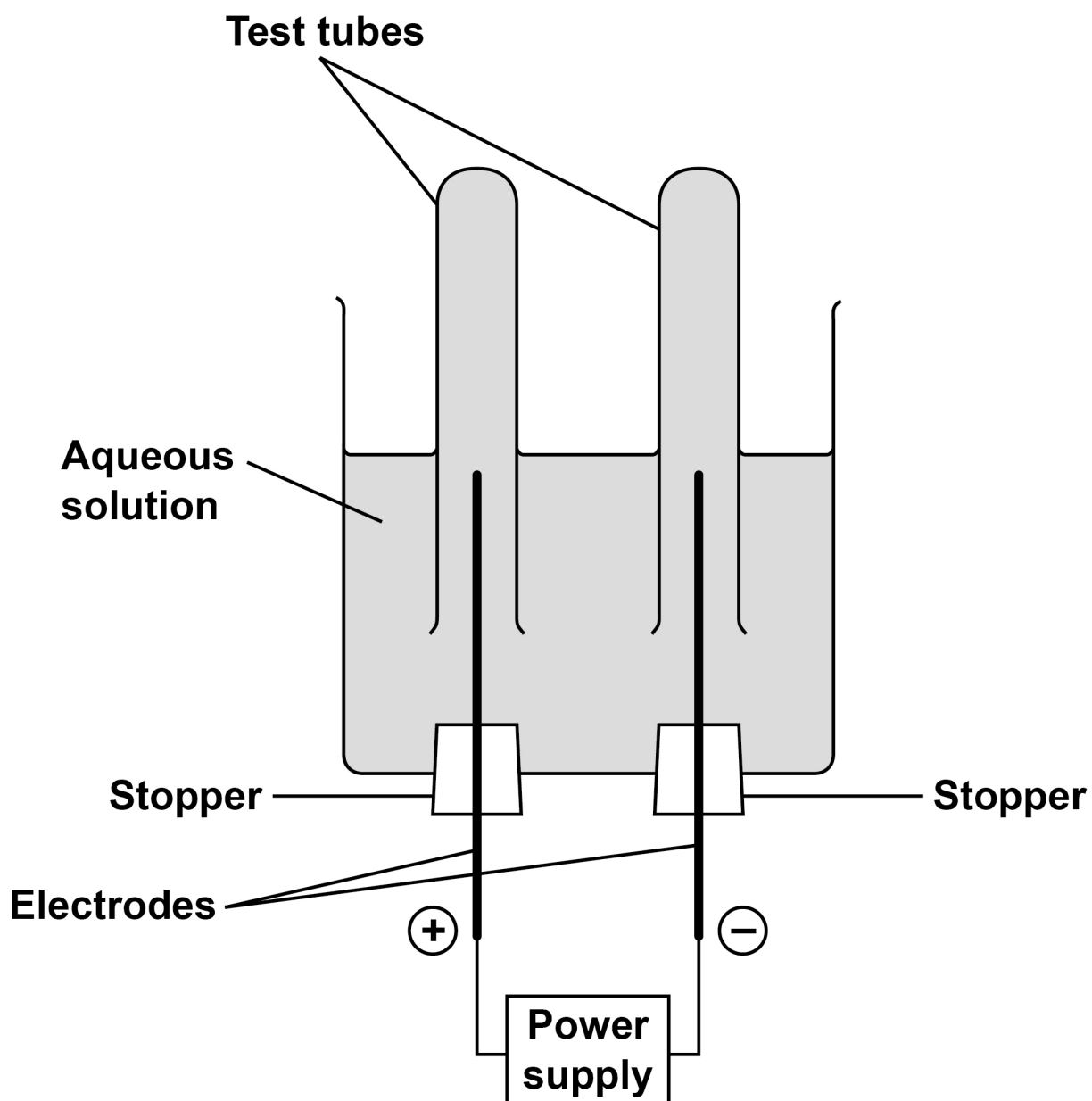
[Turn over]



A student investigated the electrolysis of an aqueous solution of a different compound.

FIGURE 7 shows the apparatus.

FIGURE 7



Hydrogen was produced at the negative electrode and oxygen was produced at the positive electrode.

07.3

Explain how oxygen was produced from water during the electrolysis of this aqueous solution. [4 marks]

[illegible]

[Turn over]



0	7	.	4
---	---	---	---

The student compared the volumes of the two gases collected.

How can the student change the apparatus in FIGURE 7, on page 44, to compare the volumes of the two gases produced more accurately?

Give ONE reason for your answer. [2 marks]

Change _____

Reason _____



0	7	.	5
---	---	---	---

The overall equation for the reaction is:



What is the volume of oxygen produced when 20 cm³ of hydrogen has been produced? [1 mark]

Tick (✓) ONE box.

☐

10 cm³

☐

20 cm³

☐

30 cm³

☐

40 cm³

[Turn over]

9



0	8
---	---

This question is about elements in the periodic table.

0	8	.	1
---	---	---	---

Argon has the atomic number 18

Explain why argon does NOT form compounds.

Answer in terms of electrons. [2 marks]

0	8	.	2
---	---	---	---

Phosphorus (P) is the element below nitrogen in the periodic table.

Predict the formula of the compound formed between phosphorus and hydrogen. [1 mark]

Formula = _____

[Turn over]



0	8	.	3
---	---	---	---

Tellurium is the element with atomic number 52

Predict whether tellurium reacts with metals.

Explain your answer.

Answer in terms of the position of tellurium in the periodic table. [2 marks]

Barium (Ba) is an element in Group 2 of the periodic table.

Barium reacts with hydrochloric acid.

08.4

Suggest TWO observations that could be made when barium reacts with hydrochloric acid. [2 marks]

1 _____

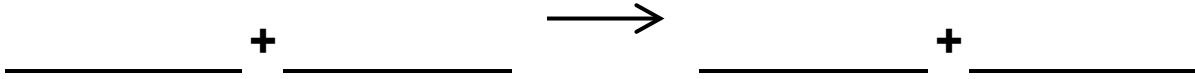
2 _____

[Turn over]



0	8	.	5
---	---	---	---

Write a balanced symbol equation for the reaction between barium and hydrochloric acid. [3 marks]



10

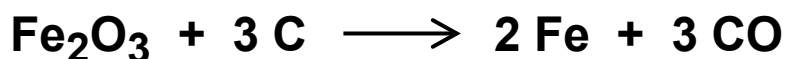


0	9
---	---

This question is about displacement reactions.

Iron is extracted from iron oxide by a displacement reaction with carbon.

The equation for the reaction is:



0	9	.	1
---	---	---	---

Which substance in the equation is reduced?

Give ONE reason for your answer.

Answer in terms of oxygen. [2 marks]

Substance reduced _____

Reason _____

[Turn over]



0	9	.	2
---	---	---	---

Which expression shows how to calculate the mass of carbon needed to produce 1 mole of iron from iron oxide?

Relative atomic mass (A_r): C = 12 [1 mark]

Tick (✓) ONE box.

<input type="checkbox"/>	$\frac{1}{3} \times 12 \text{ g}$
--------------------------	-----------------------------------

<input type="checkbox"/>	$\frac{3}{2} \times 12 \text{ g}$
--------------------------	-----------------------------------

<input type="checkbox"/>	$1 \times 12 \text{ g}$
--------------------------	-------------------------

<input type="checkbox"/>	$3 \times 12 \text{ g}$
--------------------------	-------------------------



A student investigated displacement reactions of four different metals represented by A, B, C and D.

A, B, C and D are NOT the actual chemical symbols for the metals.

The student:

- **added each metal to aqueous solutions of the metal nitrates**
- **observed whether a reaction took place.**

TABLE 6, on page 56, shows information about three of the reaction mixtures.

[Turn over]





5 6

TABLE 6

REACTION	METAL	METAL NITRATE SOLUTION	EQUATION
1	A	BNO₃	$A + 2 BNO_3 \longrightarrow 2 B + A(NO_3)_2$
2	C	A(NO₃)₂	$2 C + 3 A(NO_3)_2 \longrightarrow 3 A + 2 C(NO_3)_3$
3	C	D(NO₃)₂	no reaction



5 7

09.3

The ionic equation for REACTION 1 is:



Why is this a redox reaction? [1 mark]

Tick (✓) ONE box.

☐

A gains electrons and B^+ loses electrons.

☐

A loses electrons and B^+ gains electrons.

☐

Both A and B^+ gain electrons.

☐

Both A and B^+ lose electrons.

[Turn over]

BLANK PAGE





5 9

0 9 . 4

Which of the four metals has the greatest tendency to form positive ions?

Use TABLE 6, on page 56. [1 mark]

Tick (✓) ONE box.

☐

A

☐

B

☐

C

☐

D

[Turn over]



6 0

REPEAT OF TABLE 6

REACTION	METAL	METAL NITRATE SOLUTION	EQUATION
1	A	BNO_3	$\text{A} + 2 \text{BNO}_3 \longrightarrow 2 \text{B} + \text{A}(\text{NO}_3)_2$
2	C	$\text{A}(\text{NO}_3)_2$	$2 \text{C} + 3 \text{A}(\text{NO}_3)_2 \longrightarrow 3 \text{A} + 2 \text{C}(\text{NO}_3)_3$
3	C	$\text{D}(\text{NO}_3)_2$	no reaction



6 1

09.5

The nitrate ion has the formula NO_3^-

Which of the four metals could be aluminium?

Explain your answer.

Use TABLE 6, on the opposite page. [3 marks]

Metal _____

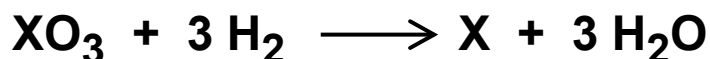
Explanation _____

[Turn over]

0	9	.	6
---	---	---	---

Metal X is extracted from an oxide of metal X by reaction with hydrogen.

The equation for the reaction is:



The percentage atom economy for obtaining metal X by this method is 77.3%.

Calculate the relative atomic mass (A_r) of metal X.

Relative atomic masses (A_r): $\text{H} = 1$ $\text{O} = 16$

[4 marks]



Relative atomic mass (A_r) = _____

[Turn over]

12



1	0
---	---

This question is about titanium dioxide (TiO_2).

1	0	.	1
---	---	---	---

Self-cleaning windows are coated with a layer of nanoparticles of titanium dioxide.

Titanium dioxide:

- helps sunlight break down dirt particles
- attracts water, so dirt is washed away by rain.

Nanoparticles of titanium dioxide are used instead of fine particles of titanium dioxide for coating self-cleaning windows.

Suggest TWO reasons why. [2 marks]

1

2



1	0	.	2
---	---	---	---

Titanium is extracted from titanium dioxide in a two-stage process.

The equation for the first stage in the process is:



Calculate the volume of chlorine gas needed to react completely with 100 kg of titanium dioxide.

Relative atomic masses (A_r): O = 16 Ti = 48

The volume of one mole of gas = 24 dm³ [6 marks]

[Turn over]





Volume = _____ dm³

8

END OF QUESTIONS



Additional page, if required.

Write the question numbers in the left-hand margin.

[illegible]

Additional page, if required.

Write the question numbers in the left-hand margin.

[illegible]

BLANK PAGE

For Examiner's Use	
Question	Mark
1	
2	
3	
4	
5	
6	
7	
8	
9	
10	
TOTAL	

Copyright information

For confidentiality purposes, all acknowledgements of third-party copyright material are published in a separate booklet. This booklet is published after each live examination series and is available for free download from www.aqa.org.uk

Permission to reproduce all copyright material has been applied for. In some cases, efforts to contact copyright-holders may have been unsuccessful and AQA will be happy to rectify any omissions of acknowledgements. If you have any queries please contact the Copyright Team.

Copyright © 2023 AQA and its licensors. All rights reserved.

WP/M/MG/Jun23/8462/1H/E4

